1. (a) For a $2p$ electron, $n = 2$, $l = 1$, $m_l = 0,\pm1$ and $m_s = \pm½$, so the possible sets of quantum numbers (n,l,m_l,m_s) are:

$$(2,1,+1,+½), (2,1,+1,–½), (2,1,0,+½), (2,1,0,–½), (2,1,–1,+½), (2,1,–1,–½)$$

(b) There are 6 possible sets of quantum numbers for each electron, so the total number of possibilities for 2 electrons is $6 \times 6 = 36$.
(c) The Pauli principle prevents the two sets from being identical. There will be 6 combinations in which the two sets are identical; eliminating these combinations leaves 30 allowed combinations.
(d) Because the n values are different, the Pauli principle does not restrict the number of combinations, to there will be 36 possible combinations.

2. (a) The two electrons in the $1s$ level have m_s of +1/2 and –1/2, so they do not contribute to the total m_s, and the same is true for the two electrons in the $2s$ level. In the $2p$ level, there are three different possible values of m_l, and for each of those values we can assign a set of quantum numbers with $m_s = +1/2$, so the maximum possible value of the total m_s is +3/2.
(b) $(n, l, m_l, m_s) = (2, 1, +1, +1/2), (2, 1, 0, +1/2), (2, 1, –1, +1/2)$
(c) There is only one possible value of the total m_l in the states that maximize m_s, and from the states listed in (b) that value is $+1 + 0 + (–1) = 0$.
(d) We could maximize the total m_l by giving the first $2p$ electron $m_l = +1$, and the second electron can also have $m_l = +1$ if we give these two electrons opposite values of m_s. The third electron cannot have $m_l = +1$, so we must assign it $m_l = 0$ and the maximum total m_l is +2.

4. (a) In beryllium ($1s^2 2s^2$) the smallest energy jump is from $2s$ to $2p$.
(b) In neon ($1s^2 2s^2 2p^6$) the smallest energy jump is from $2p$ to $3s$.
(c) From Figure 8.1 we see that the $2s \rightarrow 2p$ energy difference is smaller than the $2p \rightarrow 3s$ difference, so the minimum absorption energy would be smaller for beryllium.

8. (a) $[He]2s^2 2p^1$ (b) $[He]2s^2 2p^4$ (c) $[Ne]3s^2 3p^6$ (d) $[Ar]4s^2 3d^{10}$

11. Singly ionized lithium has two electrons. When one of those is excited to a higher level, it is screened by the one electron remaining in the 1*s* level so $Z_{\text{eff}} = 3 - 1 = 2$. The expected energy when the outer electron is excited to the 2*p* level is

$$E_n = (-13.6 \text{ eV})\frac{Z_{\text{eff}}^2}{n^2} = (-13.6 \text{ eV})\frac{2^2}{2^2} = -13.6 \text{ eV}$$

which agrees very well with the measured value of –13.4 eV. When the outer electron is in the 3*d* level, its expected energy is

$$E_n = (-13.6 \text{ eV})\frac{Z_{\text{eff}}^2}{n^2} = (-13.6 \text{ eV})\frac{2^2}{3^2} = -6.0 \text{ eV}$$

in excellent agreement with the measured value.

16. Solving Equation 8.4 for Z with $\Delta E = \dfrac{hc}{\lambda} = \dfrac{1240\ \text{eV}\cdot\text{nm}}{0.1940\ \text{nm}} = 6392\ \text{eV}$, we obtain

$$Z = 1 + \sqrt{\frac{\Delta E}{10.2\ \text{eV}}} = 1 + \sqrt{\frac{6392\ \text{eV}}{10.2\ \text{eV}}} = 26$$

so the element is iron.

17. Ca (Z = 20): $\Delta E = (10.2\ \text{eV})(Z-1)^2 = (10.2\ \text{eV})(19)^2 = 3.68\ \text{keV}$

Zr (Z = 40): $\Delta E = (10.2\ \text{eV})(Z-1)^2 = (10.2\ \text{eV})(39)^2 = 15.5\ \text{keV}$

Hg (Z = 80): $\Delta E = (10.2\ \text{eV})(Z-1)^2 = (10.2\ \text{eV})(79)^2 = 63.7\ \text{keV}$

The values computed from Moseley's law are smaller than the measured values, and the discrepancy increases as Z increases.

27. (a) For the 3*s* outer electron of sodium, inserting $E_3 = -5.14$ eV into Equation 8.1 gives

$$Z_{\text{eff}} = n\sqrt{\frac{E_n}{-13.6 \text{ eV}}} = 3\sqrt{\frac{-5.14 \text{ eV}}{-13.6 \text{ eV}}} = 1.84$$

The simple screening model predicts $Z_{\text{eff}} = 1$, so clearly the 3*s* electron is slightly penetrating the inner orbits and so is less screened by the inner electrons.
(b) For the 4*f* state,

$$Z_{\text{eff}} = n\sqrt{\frac{E_n}{-13.6 \text{ eV}}} = 4\sqrt{\frac{-0.85 \text{ eV}}{-13.6 \text{ eV}}} = 1.00$$

so the screening is complete, with the 11 positive charges in the nucleus screened by the 10 electrons in the $n = 1$ and $n = 2$ shells.

30. The wavelength difference is $\Delta\lambda = 0.59$ nm. By taking differentials of $E = hc/\lambda$, we can find the corresponding energy difference:

$$\Delta E = \frac{hc}{\lambda^2}\Delta\lambda = \frac{1240\ \text{eV}\cdot\text{nm}}{(590\ \text{nm})^2}(0.59\ \text{nm}) = 2.1\times10^{-3}\ \text{eV}$$

This energy difference comes from the interaction of a magnetic field B with a magnetic moment that we assume is of the order of 1 μ_B. The energy difference between the cases with the magnetic moment parallel to B and antiparallel to B is (see Figure 7.25) $\Delta E = 2\mu_B B$, so

$$B = \frac{\Delta E}{2\mu_B} = \frac{2.1\times10^{-3}\ \text{eV}}{2(5.8\times10^{-5}\ \text{eV/T})} = 18\ \text{T}$$

This is quite a large magnetic field, of the order of the largest that can be produced in the laboratory with superconducting electromagnets.